

# CO<sub>2</sub> Atmosphere in a jar Experiment

## Materials:

500 mL of THE PUREST DISTILLED WHITE VINEGAR (or just distilled white vinegar off the shelf)

1 tablespoon of baking soda

## Sample Table of Experiment

	(Temperature in Fahrenheit) Erlenmeyer Flask
CO <sub>2</sub> Atmosphere	105
Regular Room Air	100

## Sample Research Question

### Research Question:

Will CO<sub>2</sub> cause the temperature of a sealed flask rise faster than a sealed flask with room air that are both exposed to sunlight.

## Sample Hypothesis

### Hypothesis:

If I expose a sealed flask with extra CO<sub>2</sub> introduced to sunlight, then the temperature will rise faster than that of a sealed flask with no extra CO<sub>2</sub> because CO<sub>2</sub> is a dense gas, which allows it to absorb more heat.

NOTE Look out for seal, make sure it is completely sealed.

## Variations:

Glass vs. Plastic

## Sample table from our results

	(Temperature in Fahrenheit)
Glass	106
Plastic	105

Increase surface area

Add a reflective material as a surface (eg. aluminum foil)